

**BHARATIYA VIDYA BHAVAN'S V M PUBLIC SCHOOL,
VADODARA**

QUESTION BANK

Chapter - 1 Chemical reactions

VERY SHORT ANSWER QUESTIONS [1MARK]

1. State the law which holds good for a balanced chemical equation.
2. How do you represent physical state of reactants and products.
3. Define Oxidation and reduction.
4. Define neutralisation reaction.
5. What happens when quicklime is added to water.
6. Give one example of displacement reaction.
7. Give balanced equation when sulphur trioxide reacts with water to form sulphuric acid.
8. What are exothermic and endothermic reactions?
9. Give two ways to prevent rancidity.
10. Old silver ornaments turn black, why?

SHORT ANSWER QUESTIONS [2MARKS]

11. What is a combination reaction? Give an example.
12. Write an activity to show that iron is more reactive than copper.
13. Give one example of electrolytic and thermal decomposition.
14. Illustrate by giving suitable example the meaning of the terms oxidising and reducing agents.
15. Give methods to prevent corrosion.
16. Differentiate displacement and double displacement reactions.
17. Why does AgBr turn grey when kept in the sunlight
18. Write any two observations which may suggest that a chemical reaction has taken place. Give an example in support of your answer.
19. Give two metals which displace hydrogen from nitric acid.
20. Identify the substances oxidised, substance reduced, oxidising agent and reducing agent in the following reaction : $\text{CuO} + \text{H}_2 \longrightarrow \text{Cu} + \text{H}_2\text{O}$

21. SHORT ANSWER QUESTIONS [3MARKS]

- a. What is the colour of ferrous sulphate crystals? How does this colour change after heating?
 - b. Name the products formed on strongly heating it and write the balanced equation for the same. Name the type of reaction in this change.
22. a) Crystals of copper sulphate are heated in a test tube for sometime, What is the colour of copper sulphate crystals i) before heating ii) after heating?
 23. b) What is the source of water droplets seen on the inner sides of the test tube during heating?

24. Asha wanted her house to be white washed. She bought 30 kg of a substance 'X' and dissolved in water and observed that water started boiling without heating. Give reason for her observation, write the chemical equation involved and name the product formed.
25. What happens when aq. Solution of sodium sulphate reacts with aq. Solution of barium chloride? State the physical conditions of reactants in which reaction between them will not take place. Write the balanced chemical equation for the reaction and name the type of reaction.
26. Write the names and chemical formula of the products formed

- a) when baking soda is heated
- b) slaked lime is treated with chlorine
- c) Cu is added to AgNO_3 solution.

27. Give one example where corrosion is an advantage over disadvantage. And two examples where corrosion is disadvantage.
28. Name the products when solutions of lead nitrate and potassium iodide are mixed in a test tube. Write a balanced equation to represent this reaction and name the type of this reaction.
29. What do you mean by rancidity. How will you check that food has become rancid? Give two ways to prevent it.
30. Explain by giving three reactions that decomposition reactions are opposite of combination reactions.
31. Photosynthesis is an endothermic while respiration is an exothermic reaction, justify with equations of the reactions involved.
- LONGANSWER QUESTIONS [5MARKS]**
32. What is meant by rusting? With the help of a labelled diagram, describe an activity to find out the conditions under which iron rusts.
33. What do you mean by decomposition reactions? Explain different types of decomposition reactions with an example.